

Corrosion Inhibition of Nickel in HCl Solution by Some Indole Derivatives

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Abstract

Some indole derivatives are investigated as corrosion inhibitors for nickel in 0.5 M HCl solution using potentiodynamic polarization and electrochemical impedance spectroscopy (EIS) techniques. A significant decrease in the corrosion rate of nickel was observed in the presence of investigated indole derivatives. Potentiodynamic polarization curves revealed that these inhibitors acted as mixed-type inhibitors, affecting both cathodic and anodic corrosion processes. The adsorption of the inhibitors on nickel surface in 0.5 M HCl was found to follow Frumkin adsorption isotherm. Thermodynamic adsorption parameters (K_{ads} , ΔG_{ads}) of investigated inhibitors were calculated from the linear form of Frumkin adsorption isotherm. Activation parameters of the corrosion process were calculated and discussed. EIS was used to investigate the mechanism of corrosion inhibition. Correlation between the inhibition efficiency and the structure of these inhibitors are presented.

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Antibacterial drugs as environmentally-friendly corrosion inhibitors for carbon steel in acid medium

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Abstract

The effect of cefazolin (CZ) and cefotaxime (CT) as corrosion inhibitors for carbon steel in 0.5 M H₂SO₄ solution was investigated by use of potentiodynamic polarization, electrochemical impedance spectroscopy (EIS), electrochemical frequency modulation (EFM), and scanning electron microscopy (SEM). CZ and CT acted as mixed-type inhibitors. Inhibition increased with increasing inhibitor concentration and decreased with increasing temperature. Adsorption of the inhibitors obeyed the Langmuir adsorption isotherm. SEM confirmed inhibition by the inhibitors. Inhibition by 5 x 10⁻⁴ M CZ and 7 x 10⁻⁴ M CT approached 99.6 % and 90.9 %, respectively. The EIS and EFM results were in good agreement with the potentiodynamic data.

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Author(s): Shukla, Sudhish Kumar; Singh, Ashish Kumar; Quraishi, M. A.

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Author(s): Stoynov, Z.B.; Grafov, B.M.; Savova-Stoynova, B.; et al; Elkin, V.V.

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Role of some pyrazol-5-one derivatives as corrosion inhibitors for 316L stainless steel in 1 M HCl

Fouda, AS (Fouda, A. S.)^{1,1}; El-Ewady, GY (El-Ewady, G. Y.)^{1,1};
Fathy, S (Fathy, S.)^{1,1}

Abstract:

The effect of novel, pyrazolone derivatives, namely, (Z)-3-methyl-4-(2-m-tolylhydrazono)-1H-pyrazol-5(4H)-compound A, (Z)-4-(2-(3-methoxyphenyl)hydrazono)-3-methyl-1H-pyrazol-5(4H)-compound B, and (Z)-3-methyl-4-(2-(3-nitrophenyl) hydrazono)-1H-pyrazol-5 (4H)-compound C as corrosion inhibitors of 316L stainless steel (SS) in 1M HCl has been investigated by using weight loss, potentiodynamic polarization, electrochemical impedance spectroscopy (EIS), and electrical frequency modulation (EFM) techniques. Polarization data clearly indicated that the pyrazol-5-one derivatives behave as mixed type inhibitors. The effect of temperature on corrosion inhibition has been studied and the thermodynamic activation and adsorption parameters were calculated and discussed. EIS was used to investigate the mechanism of corrosion inhibition. EFM can be used as a rapid and nondestructive technique for corrosion rate measurements without prior knowledge of Tafel constants. The adsorption of compounds on 316L SS was found to obey Temkin adsorption isotherm.

Source: DESALINATION AND WATER TREATMENT **Volume:** 51 **Issue:** 10-12 **Pages:** 2202-2213 **DOI:** 10.1080/19443994.2012.734730 **Published:** FEB 2013

Author Keywords: 316L SS; HCl; EFM; EIS; Pyrazol-5-one derivatives

KeyWords Plus: MILD-STEEL; ACID-MEDIA; SPECTROSCOPY; TRIAZOLES

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Research Areas: Engineering; Water Resources

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Author(s): Amin, Mohammed A.; Abd El Rehim, Sayed S.; El-Naggar, M. M.; et al.

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Author(s): MCCAFFER.E; HACKERMA.N

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Author(s): Khaled, K. F.

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Author(s): Khaled, K. F.

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Author(s): ZHANG, PQ; WU, JX; ZHANG, WQ; et al.

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Published: AUG 1993

Corrosion Inhibition of Carbon Steel in 0.5 M HCl Solution Using Cationic Surfactants

Fouda, AS (Fouda, A. S.)^[1]; Elewady, YA (Elewady, Y. A.)^[1];
Abd El-Aziz, HK (Abd El-Aziz, H. K.)^[1];
Ahmed, AM (Ahmed, A. M.)^[2]

Abstract

The corrosion inhibition effect of cationic surfactants, namely: cetyl trimethyl ammonium bromide: CTAB and dodecyl trimethyl ammonium chloride: DTAC, have been used as corrosion inhibitors for C-steel in 0.5 M HCl. The inhibition efficiencies of the tested surfactants were depended on the hydrophobic chain length and the used doses of the surfactants. The results showed that the order of inhibition efficiency is CTAB > DTAC. Polarization measurements showed these surfactants are acting as mixed inhibitors for both anodic and cathodic reactions. Adsorption of these surfactants was found to follow the Langmuir's adsorption isotherm. Mixed physical and chemical adsorption mechanism is proposed. The density function theory (DFT) was used to study the structural properties of the surfactants. Inhibition efficiency values obtained from weight loss, potentiodynamic polarization, electrochemical impedance spectroscopy (IES) and electrochemical frequency modulation (EFM) are consistent.

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Author(s): Bastidas, JM; Polo, JL; Cano, E

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Curcumin Derivatives as Green Corrosion Inhibitors for alpha-Brass in Nitric Acid Solution

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Abstract

1,7-Bis-(4-hydroxy-3-methoxy-phenyl)-hepta-1,6-diene-4-arylo-3,5-dione I-V have been investigated as corrosion inhibitors for alpha-brass in 2 M nitric acid solution using weight-loss and galvanostatic polarization techniques. The efficiency of the inhibitors increases with the increase in the inhibitor concentration but decreases with a rise in temperature. The conjoint effect of the curcumin derivatives and KSCN has also been studied. The apparent activation energy (E_a) and other thermodynamic parameters for the corrosion process have also been calculated. The galvanostatic polarization data indicated that the inhibitors were of mixed-type, but the cathode is more polarized than the anode. The slopes of the cathodic and anodic Tafel lines (b_c and b_a) are maintained approximately equal for various inhibitor concentrations. However, the value of the Tafel slopes increases together as inhibitor concentration increases. The adsorption of these compounds on alpha-brass surface has been found to obey the Frumkin's adsorption isotherm. The mechanism of inhibition was discussed in the light of the chemical structure of the undertaken inhibitors.

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